
$\qquad$ Pd: $\qquad$

## Basic Stoichiometry PhET Lab rvsd 2/2011

Let's make some sandwiches!

## Introduction:

When we bake/cook something, we use a specific amount of each ingredient. Imagine if you made a batch of cookies and used way too many eggs, or not enough sugar. YUCK! In chemistry, reactions proceed with very specific recipes. The study of these recipes is stoichiometry. When the reactants are present in the correct amounts, the reaction will produce products. What happens if there are more or less of some of the reactants present?

Procedure: Go to: http://phet.colorado.edu/en/simulation/reactants-products-and-leftovers
Click the play button

## Part 1: Making Sandwiches:

## Cheese Sandwiches

1. The Cheese Sandwhich is a simulation of a two-reactant synthesis reaction. In this case, one reactant will be limiting, while the other will be in excess.
2. Take some time and familiarize yourself with the simulation.
3. Set the reaction to a simple mole ratio of $2: 1$
4. Complete the table below while making tasty cheese sandwiches (Enter values into reactants and products section at bottom for simultation)

| Bread Used | Cheese Used | Sandwiches Made | Excess Bread | Excess Cheese |
| :---: | :---: | :---: | :---: | :---: |
| 5 slices | 5 slices |  |  |  |
| 4 slices | 3 slices |  |  |  |
|  |  | 2 sandwiches | 1 slice | 0 slices |
| 6 slices |  | 3 sandwiches |  | 4 slices |

## Meat and Cheese Sandwiches

1. The Meat and Cheese Sandwhich is a simulation of a three-reactant synthesis reaction. In this case, one reactant will be limiting, while the other two will be in excess.
2. Take some time and familiarize yourself with the simulation.
3. Set the reaction to a simple mole ratio of $2: 1: 1$
4. Complete the table below while making tasty cheese sandwiches (Enter values into reactants and products section at bottom for simulation)

| Bread Used | Meat Used | Cheese Used | Sandwiches <br> Made | Excess Bread | Excess Meat | Excess Cheese |
| :---: | :---: | :---: | :---: | :---: | :---: | :---: |
| 5 slices | 5 slices | 5 slices |  |  |  |  |
| 4 slices | 3 slices | 2 slices |  |  |  |  |
|  |  |  | 2 sandwiches | 1 slice | 0 slices | 2 slices |
| 6 slices |  |  | 3 sandwiches |  | 4 slices | 1 slice |

## Part 2: Molecules:

Make Water

1. Now let's work with real chemical reaction, one that creates a very entertaining BOOM!
2. What is the mole ratio for the reaction of hydrogen and oxygen to produce water?
$\ldots \mathrm{H}_{2}+\ldots \mathrm{O}_{2} \rightarrow \ldots \mathrm{H}_{2} \mathrm{O}$
3. Complete the table below while making water $\mathrm{H}_{2} \mathrm{O}$ from hydrogen $\mathrm{H}_{2}$ and oxygen $\mathrm{O}_{2}$ :

| Hydrogen Molecules $\mathrm{H}_{2}$ | Oxygen Molecules $\mathrm{O}_{2}$ | Water Molecules $\mathrm{H}_{2} \mathrm{O}$ | Excess $\mathrm{H}_{2}$ | Excess $\mathrm{O}_{2}$ |
| :---: | :---: | :---: | :---: | :---: |
| 4 molecules | 4 molecules |  |  |  |
| 7 molecules | 6 molecules |  |  |  |
|  |  | 4 molecules | 0 molecules | 0 molecules |
| 9 moles | 8 moles |  |  |  |
|  |  | 4 moles | 1 moles | 0 moles |
| 4.0 moles | 2.5 moles |  |  |  |
| 1.5 moles |  | 1.5 moles | 0 moles | 0 moles |

4. Notice that the labels changed from molecules to moles. This does not change the mole ratio, as a mole is simply a large number of molecules. How many molecules is a mole? $\qquad$
5. Now try producing ammonia, a very important chemical in industry and farming.
6. What is the mole ratio for the production of ammonia? $\qquad$ $N_{2}+$ $\qquad$ $\mathrm{NH}_{3}$

7. Complete the table below:

| Moles $\mathrm{N}_{2}$ | Moles $\mathrm{H}_{2}$ | Moles $\mathrm{NH}_{3}$ | Excess $\mathrm{N}_{2}$ | Excess $\mathrm{H}_{2}$ |
| :---: | :---: | :---: | :---: | :---: |
| 3 moles | 6 moles |  |  |  |
| 6 moles | 3 moles |  |  |  |
|  |  | 4 moles | 2 moles | 0 moles |
| 1.5 moles | 4.0 moles |  |  |  |

8. Combustion of hydrocarbons like methane $\mathrm{CH}_{4}$ produce two products, water and carbon dioxide $\mathrm{CO}_{2}$.
9. What is the mole ratio for the combustion of methane? _ $\mathrm{CH}_{4}+\ldots \mathrm{O}_{2} \rightarrow \ldots \mathrm{CO}_{2}+\ldots \mathrm{H}_{2} \mathrm{O}$
10. Complete the table below: WATCH FOR FRACTIONS

| mol CH4 | $\mathrm{mol} \mathrm{O}_{2}$ | $\mathrm{mol} \mathrm{CO}_{2}$ | $\mathrm{mol} \mathrm{H}_{2} \mathrm{O}$ | Excess mol CH ${ }_{4}$ | Excess mol $\mathrm{O}_{2}$ |
| :---: | :---: | :---: | :---: | :---: | :---: |
| 4 mol | 4 mol |  |  |  |  |
| 3 mol | 6 mol |  |  |  |  |
|  |  | 2 mol | 4 mol |  |  |
|  |  | 3 mol |  | 0 mol |  |
|  |  |  |  | 2 mol | 0 mol |
|  |  |  |  | 3 mol | 1 mol |

## Part 3: Game:

Game: Level 1
Fill in the chart and include the correct formulas, no leftovers. Play for time after first time.

|  | Reactants | $\rightarrow$ | Products | Best time |
| :--- | :--- | :--- | :--- | :--- |
| \#1 |  |  |  |  |
| \#2 |  |  |  |  |
| \#3 |  |  |  |  |
| \#4 |  |  |  |  |
| \#5 |  |  |  |  |

## Game: Level 2

Fill in the chart and include the correct formulas, no leftovers. Play for time after first time.

|  | Reactants | $\rightarrow$ | Products | Best time |
| :--- | :--- | :--- | :--- | :--- |
| $\# 1$ |  |  |  |  |
| $\# 2$ |  |  |  |  |
| $\# 3$ |  |  |  |  |
| $\# 4$ |  |  |  |  |
| $\# 5$ |  |  |  |  |

## Game: Level 3

Fill in the chart and include the correct formulas, no leftovers. Play for time after first time.

|  | Reactants | $\rightarrow$ | Products | Best time |
| :--- | :--- | :--- | :--- | :--- |
| \#1 |  |  |  |  |
| \#2 |  |  |  |  |
| \#3 |  |  |  |  |
| \#4 |  |  |  |  |
| \#5 |  |  |  |  |

